

Answer on Question #50372, Chemistry, Other

What is the pH of a solution with a  $[H_3O^+]$  concentration of  $1.5 \times 10^{-13}$  M?

**Solution:**

$$[H^+] = [H_3O^+] = 1.5 \times 10^{-13}$$

$$pH = -\log_{10}[H^+]$$

$$pH = -\log_{10} 1.5 \times 10^{-13} = 12.82$$

**Answer:**

**pH = 12.82**

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