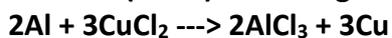


Answer on Question #50325, Chemistry, Other

Task:

How many grams of aluminum (Al) would react completely with 1350 grams of copper (II) chloride (CuCl₂) according to the following equation?



Answer:

$$v = \frac{m}{M}$$

$$M(\text{Al}) = 27 \text{ g/mol}$$

$$M(\text{CuCl}_2) = 134.5 \text{ g/mol}$$

$$v(\text{CuCl}_2) = \frac{1350}{134.5} = 10.04 \text{ mol}$$

$$v(\text{Al}) = \frac{2 \cdot v(\text{CuCl}_2)}{3} = 6.7 \text{ mol}$$

$$m(\text{Al}) = v(\text{Al}) \cdot M(\text{Al})$$

$$m(\text{Al}) = 6.7 \cdot 27 = 180.7 \text{ g}$$