

Answer on Question #50210 - Chemistry – Other

Question

What is the mass of 5.0 moles of $\text{Ba}(\text{CN})_2$?

Answer:

Mass of the substance equals:

$$m = nM$$

Molar mass of $\text{Ba}(\text{CN})_2$ is 189 g/mol. Therefore:

$$m(\text{Ba}(\text{CN})_2) = n(\text{Ba}(\text{CN})_2) \cdot M(\text{Ba}(\text{CN})_2) = 5.0 \cdot 189 = 945 \text{ g}$$

Answer: 945 g