

## Answer on Question#50205 - Chemistry, Other

**Task:** How many moles does 80.0 grams of H<sub>2</sub>O represent?

**Solution:**

1. Find the molecular mass (M) of H<sub>2</sub>O. It is the mass of 1 mole of H<sub>2</sub>O. Atomic mass ( $a_m$ ) of each chemical element we can find at the Periodic table.

$$M[H_2O] = 2 \times a_m [H] + a_m [O] = 2 \times 1 + 16 = 18$$

2. Count the amount of moles ( $v$ ) in 80.0 grams of H<sub>2</sub>O:

$$v = 80.0 \div 18 = 4, (4) \text{ moles}$$

**Answer:** 4,(4) moles.

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