

Answer on Question #49815 – Chemistry - Inorganic Chemistry

How much CuCl_2 is needed to prepare 325 g of 1.00% (w/w) solution?

Solution:

$$m_{\text{sol}} = 325 \text{ g}$$

$$w(\text{CuCl}_2) = 1.0 \%$$

$$m(\text{CuCl}_2) = ?$$

$$w(\text{CuCl}_2) = \frac{m(\text{CuCl}_2)}{m_{\text{sol}}} \times 100\%$$
$$m(\text{CuCl}_2) = \frac{w(\text{CuCl}_2) \times m_{\text{sol}}}{100\%}$$
$$m(\text{CuCl}_2) = \frac{1.00\% \times 325 \text{ g}}{100\%} = 3.25 \text{ g}$$

Answer: 3.25 g of CuCl_2 is needed.