How much $CuCl_2$ is needed to prepare 325 g of 1.00% (w/w) solution?

Solution:

m_{sol}=325 g

$w(CuCl_2)$=1.0 %

$m(CuCl_2)$-?

$$w(CuCl_2) = \frac{m(CuCl_2)}{m_{sol}} \times 100\%$$
$$m(CuCl_2) = \frac{w(CuCl_2) \times m_{sol}}{100\%}$$
$$m(CuCl_2) = \frac{1.00\% \times 325\ g}{100\%} = 3.25\ g$$

Answer: 3.25 g of $CuCl_2$ is needed.