

Answer on Question #49332, Chemistry, Organic Chemistry

Question: Assign oxidation states to each atom in each of the following species.

Part B

Fe^{2+}

Express your answer as a signed integer.

the oxidation state of Fe =

Part C

CuCl_2

Express your answer as signed integers separated by a comma

the oxidation states of Cu, Cl =

Part D

CO

Express your answer as signed integers separated by a comma.

The oxidation states of C, O =

Part E

$\text{Cr}_2\text{O}_7^{2-}$

Express your answer as signed integers separated by a comma.

The oxidation states of Cr, O =

Part F

HCO_3^-

Express your answer as signed integers separated by commas.

the oxidation states of H, C, O =

Answer:

1. $\text{Fe}^{2+} = +2$
2. $\text{Cu, Cl}_2 = +2, -1$
3. $\text{C, O} = +2, -2$
4. $\text{Cr, O} = +6, -2$
5. $\text{H, C, O} = +1, +4, -2$