

Answer on the question #49329, Chemistry, Other

Question:

using word equations and turning them into chemical equations and balancing them. (also writing the state of matter)

- 1) a double displacement reaction occurs between aqueous calcium hydroxide and lithium fluoride.
- 2) Aqueous forms of lead II nitrate and potassium sulfate react, with a product being a precipitate of lead II sulfate.
3. Aqueous sodium carbonate and silver I nitrate are reactants in a double displacement reaction.
4. when crystalline $C_6H_{12}O_6$ is burned into oxygen a combustion reaction occurs.

Answer:

- 1) Calcium hydroxide(aq) + lithium fluoride(aq) = calcium fluoride + lithium hydroxide(aq)
 $Ca(OH)_2(aq) + 2LiF(s) = CaF_2(aq) + 2LiOH(aq)$
- 2) Lead nitrate (aq) + sodium sulfate (aq) = lead sulfate (s) + potassium nitrate (aq)
 $Pb(NO_3)_2(aq) + K_2SO_4(aq) = PbSO_4(s) + 2KNO_3(aq)$
- 3) Sodium carbonate (aq) + silver nitrate (aq) = silver carbonate (s) + sodium nitrate (aq)
 $Na_2CO_3(aq) + 2AgNO_3(aq) = 2Ag_2CO_3(s) + NaNO_3(aq)$
- 4) Glucose (cris) + oxygen(g) = carbon monoxide (g) + water (g)
 $C_6H_{12}O_6(crist) + O_2(g) = CO_2(g) + H_2O(g)$