

Answer on Question #49144 - Chemistry – Other

Question

A 100 ml solution contains 4.7g MgCl₂ (MM=94).The density of the solution is 1.03g/ml.

Express the concentration of the solution in: c) N

Answer:

Equivalent molar mass of MgCl is:

$$M_{eq} = M / 2 = 94 / 2 = 47 \text{ g-eq/mol}$$

Normal concentration of magnesium chloride is:

$$C_N = \frac{m}{M_{eq}V} = \frac{4.7 \cdot 1000}{47 \cdot 100} = 1 \text{ N}$$

Answer: 1 N