## Answer on Question \#49142, Chemistry, Other

Question: A 100 ml solution contains 4.7 g MgCl ( $\mathrm{MM}=94$ ). The density of the solution is $1.03 \mathrm{~g} / \mathrm{ml}$.Express the consentration of the solution in: a) \% by mass

Answer: The mass of the solution is $100 \mathrm{ml}{ }^{*} 1.03 \mathrm{~g} / \mathrm{ml}=103 \mathrm{~g}$. If 103 g of solution contains 4.7 g of MgCl 2 then the mass of MgCl 2 in 100 g of solution will be $\left(100^{*} 4.7\right) / 103=4.56 \mathrm{~g}$.

So, the \% of MgCl 2 by mass is $4.56 \%$.

