

Answer on Question #48803, Chemistry, Other

**Task:**

Calculate the  $\text{H}_2\text{C}_2\text{O}_4$  (aq) molarity if 10.00 ml of its solution reacts completely with 31.5 ml of a 0.0158 M  $\text{KMnO}_4$ ?

**Answer:**

$$C_M(\text{H}_2\text{C}_2\text{O}_4) \cdot V(\text{H}_2\text{C}_2\text{O}_4) = C_M(\text{KMnO}_4) \cdot V(\text{KMnO}_4)$$

$$C_M(\text{H}_2\text{C}_2\text{O}_4) = \frac{0,0158 \cdot 31,5}{10,00} = 0,05 \text{ M}$$