Question #47870, Chemistry, Other

How many grams of NaCl (s) are needed to make 2.5 liters of 0.75 M solution of Nacl?

Answer:

C(NaCl)=0.75 M = 0.75 mol/L

M(NaCl)=58.4 g/mol

$$n(NaCl)=C*V$$

$$n=m/M$$

$$m/M=C*V$$

$$m=C*V*M$$

$$m(NaCl)=0.75*58.4*2.5=\mathbf{109.5\ (g)}$$