Answer on Question #47786 - Chemistry – Other

Question

Number of moles present in 22 g of 2 CO gas

Answer:

Molar mass of CO gas equals:

$$M(CO) = M(C) + M(O) = 12.011 + 15.999 = 28.01 \; g/mol$$

Number of moles in 22 g of CO gas is:

$$n(CO) = \frac{m(CO)}{M(CO)} = \frac{22}{28.01} = 0.785 \; mol$$

Answer: 0.785 mol