Answer on the question #47453, Chemistry, Other

Question:

Mass of test tube & 250mL beaker 130.729g. Mass of test tube compound A 131.756g. Heat test tube and compound A. Mass of test tube 250 mL beaker and compound B 131.369g Add 2mL concentrated nitric acid to test tube and heat. Add 2mL of HNO3 and heat. Mass of test tube 250mL beaker compound C (CuO) 131.071g. Mass of test tube 250mL beaker and copper 130.997g.

Using your experimental data calculate: mass of copper in compounds A,B,C. Mass of other element in compound A and B. Mass of oxygen in compound C.

Answer:

Mass of the compound A:

131.75 – 130.729 = 1.46 g

Mass of compound B: 131.369 – 130.729 = 0.64 g

Mass of compound C: 131.071 – 130.729 = 0.342 g

Mass of cooper: 130.997 – 130.729 = 0.268 g

Mass of other compound: in A: 1.46 – 0.268 = 1.192 g

In B: 0.64 – 0.268 = 0.372 g

Mass of oxygen in C: 0.342 – 0.268 = 0.074 g