

Question#47337 – Chemistry – Other

Question:

if 12.56g of H_3PO_4 was produced what is the percent yield

Answer:

If the theoretical mass of the H_3PO_4 is m_{theor} (calculated according to the equation reaction), then the percent yield is:

$$\omega (\text{H}_3\text{PO}_4) = 12.56 \text{ g} / m_{\text{theor}} \times 100\%$$