

Question #45933, Chemistry, Other

If lithium and fluorine react, which has more attraction for an electron? Why?
In a bond between fluorine and iodine, which has more attraction for an electron? Why?

Answer:

The oxidation of lithium fluoride, fluorine is more attractive to electrons. Cause a fluorine atom is more electronegative than lithium. Nuclear charge increases in the period, and the atomic radius decreases. The atomic radius of fluorine 150 pm, and the atomic radius of lithium 220 pm. Therefore, the compound called lithium fluoride, but not vice versa.

In bond between iodine and fluorine, fluorine is also more attractive to the electron. Fluoro less and more electronegative than iodine. Therefore, fluorine pulls the electron belonging to iodine. The atomic radius of iodine 198 pm. Electronegativity iodine - 1.9 and fluorine - 4.0.