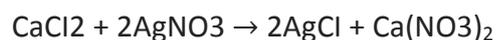**Answer on Question #45772, Chemistry, Inorganic Chemistry**

Question: How many g of caci2 does it take to produce 14.3 g of agcl when treated with excess agno3? ca(no3)2 is the other product

Answer:

An equation of this reaction is:

$CaCl2 + 2AgNO3 \rightarrow 2AgCl + Ca(NO3)_2$

The mass of one mole of AgCl = 108 + 35.5 = 143.5

The mole of AgCl = 14.3g/143.5 = 0.1 mole

n(CaCl2)/n(AgCl) = ½, so n(CaCl2) = n(AgCl)/2 = 0.1/2 = 0.05moles

The mass of 1 mole of CaCl2 = 40+35.5*2 = 111

So, m(CaCl2) = 0.05moles*111 = 5.55g