

### Question #45455, Chemistry, Physical Chemistry

DISSOLUTION OF NA<sub>2</sub>CO<sub>3</sub> IN WATER IS A PHYSICAL CHANGE NOT A CHEMICAL CHANGE . WHY?

**Answer:**

Dissolving the salt in water is difficult to call a chemical change as Na<sup>+</sup> and Cl<sup>-</sup> form a strong base and strong acid and therefore is not hydrolyzed. If you look at the definition of a chemical reaction and there is written that "the passage of a chemical reaction old bonds are broken and formed new, electrons are redistributed." Here new bonds are formed, and the process is similar to the bursting of two magnets because that is dominated by ionic bonds. Therefore, most chemists say that it is a physical process.