Answer on Question #45030, Chemistry, Inorganic Chemistry

$$K_{sp}\,(BaSO_4) = [Ba^{2+}] \cdot [SO_4{}^{2-}]$$

$$[Ba^{2+}] = [SO_4{}^{2-}]$$

$$K_{sp}\,(BaSO_4) = [Ba^{2+}]^2 = (3.9 \cdot 10^{-5})^2 = 1.5 \cdot 10^{-9}$$

Answer E