

Answer on Question #45026 – Chemistry – Other

Question

Which value best determines whether a reaction is spontaneous?

- A) change in Gibbs free energy, ΔG
- B) change in entropy, ΔS
- C) change in kinetic energy, ΔKE
- D) change in enthalpy, ΔH
- E) change in heat of formation, ΔH_{of}

Answer

Correct option is **A) change in Gibbs free energy, ΔG**

Reaction is spontaneous if $\Delta G < 0$, and is not spontaneous if $\Delta G > 0$.