

## Answer on Question #44464 – Chemistry – Other

### Question

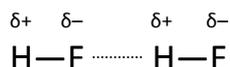
What are the attractive forces between the molecules of HF and HCl?

### Answer

Both HF and HCl molecules are very polar: dipole moment of HF = 1.86 D and dipole moment of HCl = 1.05 D, so HF and HCl form permanent dipoles:



So, the attractive forces between the molecules of HF and HCl are dipole-dipole interaction.



The hydrogen bond is often described as a strong electrostatic dipole-dipole interaction. The attractive force between molecules of **HF** is **hydrogen bonding**. The interaction between molecules of **HCl** is not strong enough to be called hydrogen bonding, so it is just **dipole-dipole interaction**.