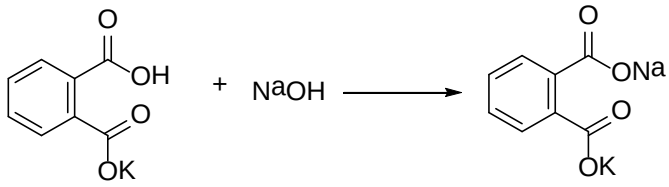
Answer on Question #42974, Chemistry, Other

Question:

1.4 g sample of KHP required 24.11 cm^3 of NaOH for its complete neutralization. Calculate the molarity of the NaOH used in titration?

Solution:



1.4 g of KHP is equal to
n(KHP) = 1.4 g/204.2 g*mol^{-1} = 6.86 mmol
n(KHP) = n(NaOH)
C(NaOH) = 0.00686 mol/0.02411 L = 0.284 mol/L

Answer: 0.284 mol/L