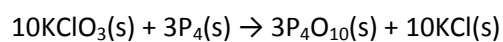


1. The reaction is:



2. Convert to moles:

$$n_{\text{KClO}_3} = \frac{m_{\text{KClO}_3}}{M_{\text{KClO}_3}} = \frac{52.9}{122.55} = 0.432 \text{ (mole)}$$

3. Multiply by mole ratio:

$$n_{\text{P}_4\text{O}_{10}} = n_{\text{KClO}_3} \cdot \frac{3}{10} = 0.432 \cdot 0.3 = 0.129 \text{ (mole)}$$

4. Convert to grams:

$$m_{\text{P}_4\text{O}_{10}} = n_{\text{P}_4\text{O}_{10}} \cdot M_{\text{P}_4\text{O}_{10}} = 0.129 \cdot 283.89 = 36.62 \text{ (g)}$$

Answer: 36.62 g.