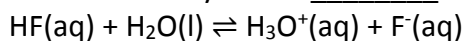


Answer on Question #41600, Chemistry, Other

Task:

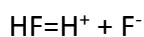
The equilibrium constant for the reaction below has the value $K_a=3.5\times 10^{-4}$. In this reaction the Brønsted-Lowry acid is _____.



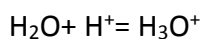
Express your answer as a chemical formula.

Answer:

A Bronsted-Lowry acid is defined as anything that releases H^+ ions. In this case a Fluorine acid dissociates to:



At the same time H_2O molecule accepts the H^+ proton.



So if the HF releases the H^+ , it is the Brønsted-Lowry acid in this reaction.