

Answer on question #40553, Chemistry, Other

Question:

If you dissolve 25.5 g KBr in enough water to make 1.75 L of solution, what is the molarity of the solution?

Solution:

We need to know the amount of KBr:

$$N(\text{KBr}) (\text{mol}) = 25.5 / 119.0 = 0.21$$

Now we can find the molarity of the solution:

$$M(\text{KBr}) (\text{mol/L}) = 0.21 / 1.75 = 0.12$$

Answer: **0.12**