

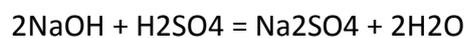
## Answer on Question#39838, Chemistry, Other

### Question:

How many milliliters of 0.121 M H<sub>2</sub>SO<sub>4</sub> are needed to neutralize 0.150g of NaOH?

### Answer:

An equation of the reaction is:



Moles of NaOH =  $0.150/40 = 0.00375$  moles.

The mole ratio of this equation is:

$$n(\text{NaOH})/n(\text{H}_2\text{SO}_4) = 2/1$$

The moles of H<sub>2</sub>SO<sub>4</sub> =  $0.00375/2 = 0.001875$  moles

The volume of H<sub>2</sub>SO<sub>4</sub> we can be found from formula:

$$C = n/V$$

$$V = 0.001875/0,121 = 0.0155\text{L} = 15.5\text{ml}$$