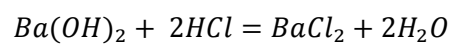


Answer on Question#39837 - Chemistry - Other

Question:

How many milliliters of 0.100 M HCl are needed to completely neutralize 52.0mL of 0.110 M Ba(OH)₂ solution?

Solution:



$$n(\text{Ba(OH)}_2) \text{ (mmole)} = 52.0 \times 0.110 = 5.72;$$

$$n(\text{HCl}) \text{ (mmole)} = 2 \times 5.72 = 11.44;$$

$$V(\text{HCl}) \text{ (ml)} = 11.44 / 0.100 = 114.4;$$

Answer: 114.4