Answer on Question#39836 - Chemistry - Other

Question:

A volume of 20.0mL of a 0.590M HNO3 solution is titrated with 0.900M KOH. Calculate the volume of KOH required to reach the equivalence point.

Solution:

$C_1V_1 = C_2V_2$

20 ml * 0.590M = 0.900 M * V_2

V_2 = (20 ml * 0.590M) / 0.900 M = 13,11 ml

Answer: 13.11 ml