

### Answer on Question#37568 - Chemistry - Inorganic Chemistry

$$C ( AgNO_3 ) = 0.01 \text{ mol} \cdot \text{dm}^{-3}$$

$$V (\text{sol. } AgNO_3) = 100 \text{ cm}^3 = 0.1 \text{ dm}^3$$

$$M ( AgNO_3 ) = 169.87 \text{ g} \cdot \text{mol}^{-1}$$

Determine the mass of silver nitrate:

$$m ( AgNO_3 ) = M * n = M * (C * V) = 169.87 * 0.01 * 0.1 = 0.16987 \text{ g}$$

The calculated mass of salt placed in a beaker, and then added water. Bring the volume up to 100 ml then stir thoroughly until complete dissolution of silver nitrate. For better dissolution water temperature should not be less than 20 degrees Celsius.

**Answer:** 0.16987g (  $AgNO_3$  ) dissolved in 100 ml (  $H_2O$  )