We have such equation for titration:

$N_1V_1 = N_2V_2$

V is the consumed volume of solution, N is the solution normality:

N = c/z

where z is change in oxidation state for red-ox titrations, amount of available proton or hydroxide ions for acid-base titrations and so on.

$\mathrm{c} = \mathrm{n}/\mathrm{V} = \frac{m/M}{V}$, hence $\mathrm{N} = \mathrm{c}/\mathrm{z} = \frac{m}{M/z \cdot V}$

$$\frac{m}{V} = \frac{N_2 V_2}{V_1} \cdot \frac{M}{z}$$

The value m/V can be called alkalinity of the solution in grams. So your value seemed to be the equivalent molar mass of your alkali M/z.