How would you prepare a solution S of pottasium iodide with a mass concentratiin of 8.4g.l-1 in v=250ml? What is the molar concentration C of the solution?

Solution:

m (KJ)(g) = 8.4*250/1000 = 2.1

$$C = \frac{n}{V} = \frac{m}{MV}$$

C (mole/l) = 2.1/(166*0.25) = 0.0506

Answer: 2.1 g of KJ; C = 0.0506 M