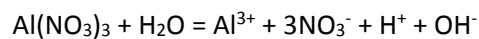


Task:

what happens when aluminium nitrate reacts with water?

Answer:

In water $\text{Al}(\text{NO}_3)_3$ dissociates into ions:



But Al^{3+} with OH^- ions tends to form a precipitate $\text{Al}(\text{OH})_3$. But we have strong acid in this solution and this doesn't allow to form aluminium hydroxide, because HNO_3 reacts with $\text{Al}(\text{OH})_3$ to form aluminium nitrate.

That's why we have in the solution

