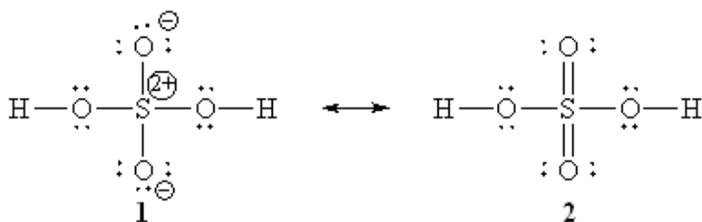


Lewis diagrams are to be used to examine mechanisms so that knowing which parts of a molecule are electron deficient and which are electron rich are vital. It is ideal to have a formal charge of 0 for as many of the atoms as possible. If a formal charge of 1- is located next to a formal charge of 1+, the formal charges can be minimized by having a lone pair of electrons, located on the atom with the 1- charge become a bonding pair of electrons that is shared with the atom that has the 1+ formal charge (this can be visualized in the same way as the formation of multiple bonds were above).



Structures 1 and 2 are resonance structures of each other, but structure 2 is the lower energy structure, even though it violates the octet rule. Sulfur can accommodate more than eight electrons, and the formal charges in structure 2 are all zero.