A graduated cylinder contains 23.0mL {\rm mL} of water. What is the new water level after 36.2g {\rm g} of silver metal with a density of 10.5 g/mL {\rm g/mL} is submerged in the water?

Solution

d = m/V; V = m/d; whrere: d – density, m – mass, V – volume.

So, V = 36.2/10.5 = 3.45 mL of silver

A new water level will be 23+3.45 = 26.45 mL