Reaction for this process is next:

$$
\mathrm{C}_{2} \mathrm{H}_{5} \mathrm{OH}+3 \mathrm{O}_{2}=2 \mathrm{CO}_{2}+3 \mathrm{H}_{2} \mathrm{O}
$$

As can you see the molar ratio is $1: 3\left(\mathrm{C}_{2} \mathrm{H}_{5} \mathrm{OH}: \mathrm{O}_{2}\right)$
So if you have 15 mol of ethanol, you need $15 * 3=45.0 \mathrm{~mol}$ of $\mathrm{O}_{2}$ to completely burn all your alcohol.

