If you know density and volume is given too, you can find mass of substance. It is the first thing you should do in the task like this:

$m = d * V$

$m = 1.54*575 = 885.5$ g

Now, you need to use Avogadro constant. In chemistry and physics, the Avogadro constant (symbols: L, NA) is defined as the number of constituent particles (usually atoms or molecules) in one mole of a given substance. It has dimensions of reciprocal mol and its value is equal to $6.022*10^{23}$ mol^{-1}.

As you can see, you also need amount:

$n = m/Mw$, where Mw is molecular weight, for glucose it is 180.

$n = 885.5 / 180 = 4{,}92$ mol

So, $N = NA * n$ (N is number of molecules)

$N = 6.022*10^{23} * 4.92 =$ **$2.962*10^{24}$ molecules**