The equation for this reaction is next:

$4 NH_3 + 5 O_2 \rightarrow 4 NO + 6 H_2O$
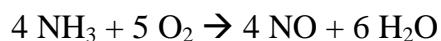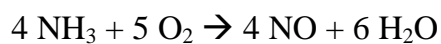
As you can see volume and mole ratio between NH_3 and O_2 is 4:5.

For example 40 L of NH_3 reacts with 50 L of O_2. In your case, where volume of ammonia is 60 L and volume of oxygen is 50L, the last one is definitely in shortage, so it is limiting reactant.

If you have 60 L of ammonia you need X L of oxygen for its complete reaction:

  60 L   x

$4 NH_3 + 5 O_2 \rightarrow 4 NO + 6 H_2O$

  4   5

x= 60*5/4 =75 L but you have only 50 L, oxygen is limiting agent.