Scheme for this reaction is:

$CxHy + n\ O_2 \rightarrow xCO_2 + 2/y\ H_2O$

But we actually don't need it, because aim of the task is to calculate % of C and H. It is logical that all C that forms CO_2 was component of CxHy (organic compound), the same with H from water.

For finding % we need to find amounts of H and C first.

$nC = nCO_2$

$nH = 1/2n\ H_2O$

$nC=nCO_2 = 1.32/44 = 0,03$ mol

nH = ½ * 0.315/18 =0,00875 mol

Percent of C =0,03 / (0,03+0,00875) *100% =77.42 %

Percent of H = 22.58%