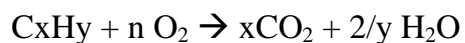


Scheme for this reaction is:



But we actually don't need it, because aim of the task is to calculate % of C and H. It is logical that all C that forms CO_2 was component of C_xH_y (organic compound), the same with H from water.

For finding % we need to find amounts of H and C first.

$$n\text{C} = n\text{CO}_2$$

$$n\text{H} = 1/2n \text{H}_2\text{O}$$

$$n\text{C}=n\text{CO}_2 = 1.32/44 = 0,03 \text{ mol}$$

$$n\text{H} = 1/2 * 0.315/18 = 0,00875 \text{ mol}$$

$$\text{Percent of C} = 0,03 / (0,03+0,00875) * 100\% = 77.42 \%$$

$$\text{Percent of H} = 22.58\%$$