

The scheme of conversation KClO_3 into AgCl is:



Mole ratio of KClO_3 and AgCl is 1: 1

Amount of AgCl is the same as KClO_3 and it is:

$$n = 185 / 143 = 1,2934 \text{ mol}$$

$$\text{Mass of } \text{KClO}_3 \text{ is } M_w \text{ of it } * \text{ amount} = 122.5 * 1,2934 = 158,48 \text{ g}$$

$$\% \text{ of } \text{KClO}_3 \text{ is } 158,48/350 * 100\% = 45,28\%$$