The chemical equation for this process is next:

$$CaCO_3 + 2\,HCl \rightarrow CaCl_2 + CO_2 + H_2O$$

As you can see, the molar ratio between calcium carbonate and calcium chloride is 1: 1, it means that one mole of calcium chloride must react with one moles of hydrochloric acid.

Theoretical amount of $CaCO_3$ is m/Mw =50 / 100 = 0.5 mol

Real amount of $CaCO_3$ is equal to calcium chloride amount and it is : 10 / 111 = 0,090 mol

Yield of reaction is 0.09/0.5 * 100% =18,0 %