Equation $N_1V_1=N_2V_2$ can help to solve this task. N is normality, Normality (N) is another ratio that relates the amount of solute to the total volume of solution.

It is defined as the number of equivalents per liter of solution:

Normality = number of equivalents / 1 L of solution

There is a very simple relationship between normality and molarity:

$N = n \times M$ (where n is an integer)

n for H_2SO_4 is 2 ($2H^+$)

So $N_{acid} = 2M_{acid}$

n for NaOH is 1 N=M

If $V_{acid} = 1$ ml

$V_{base} = 12.6$ ml

$N_{base} = 0.1000$

N of acid can be found as:

$N_{acid} = N_{base} V_{base} / V_{acid} = 0.1000 * 12.6 / 1 = 1.26$

Molarity is : N/2 = 1.26 / 2 = **0.63 M**