In chemistry, the molar concentration, c_i is defined as the amount of a constituent n_i (usually measured in moles – hence the name) divided by the volume of the mixture V:

$$c_i = \frac{n_i}{V}$$

It is also called molarity, amount-of-substance concentration, amount concentration, substance concentration, or simply concentration. The volume V in the definition $c_i = n_i/V$ refers to the volume of the solution, *not* the volume of the solvent. One liter of a solution usually contains either slightly more or slightly less than 1 liter of solvent because the process of dissolution causes volume of liquid to increase or decrease. So, if you have one liter of commercial acid the weight of it is:

m = 1000 ml * 1.787 g/ml = 1767 g

The mass of acid in this case is:

m_a = 1767 * 86% / 100% = 1519,62 g

Amount of acid is:

1519,62 / 98 = **15,5** mol the same is molarity (it was calculated for one liter)

Second part:

If you need 0,2 M solution, it means that it is 0,2 mol in one liter.

So you need dissolve some volume X of 15,5 M solution that includes 0,2 mol of acid.

If 1000 ml includes 15.5 mol

X ml includes 0,2 moles

x = 12.9 ml

So you need mix 12.9 ml of commercial acid with 987,1 ml of water. (Together it's one liter.)