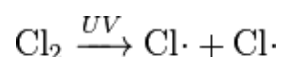In chemical equations, free radicals are frequently denoted by a dot placed immediately to the right of the atomic symbol or molecular formula as follows:

$$Cl_2 \xrightarrow{UV} Cl\cdot + Cl\cdot$$

Chlorine gas can be broken down by ultraviolet light to form atomic chlorine radicals.

It was step of Initiation. Initiation reactions are those that result in a net increase in the number of free radicals. They may involve the formation of free radicals from stable species as in Reaction 1 above or they may involve reactions of free radicals with stable species to form more free radicals.

Next step: Propagation reactions are those reactions involving free radicals in which the total number of free radicals remains the same. I your case it is :

$(CH_3)_3C\text{-}H + \cdot Cl \rightarrow (CH3)3C\cdot + HCl$
$(CH3)3C\cdot + Cl\text{-}Cl \rightarrow \textbf{\textcolor{red}{(CH}_3\textbf{)}_3\textbf{C-Cl}} + \cdot Cl$