

Buffer is composed of a weak acid and its conjugate base, or a weak base and its conjugate acid, but since HCl and H₂SO₄ are strong acids - they would not make a buffer solution.

pH for this mixture is equal to $-\log[\text{H}^+]$, concentration of $[\text{H}^+]$ is equal to $[\text{HCl}] + 2[\text{H}_2\text{SO}_4]$ (because last one is strong diprotic acid),so

$$[\text{H}^+] = 0.1 + 2 \cdot 0.1 = 0.3\text{M}$$

$$\text{pH} = -\log(0.3) = 0.52$$