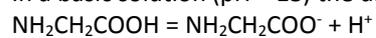


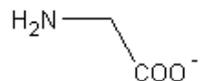
the dominant forms of glycine $\text{NH}_2\text{CH}_2\text{COOH}$ in basic solution $\text{pH}=13$ and in acidic solution $\text{pH}=3$ are, respectively?

Solution:

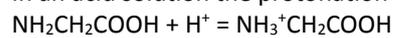
In a basic solution ($\text{pH} = 13$) the dissociation of acid group occurs:



The dominant form is



In an acid solution the protonation of amine group takes place:



The predominant form is

