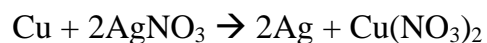


The equation for this reaction is:



The moles ratio of Cu: Ag is 1:2 (from equation).

Amount of Cu is m/M_w , where M_w is molecular mass,

$$n = m / M_w = 5.00 / 64 = 0.077 \text{ mol}$$

So in the end of reaction amount of Ag must be $2 \times 0.077 = 0.154 \text{ mol}$

$$m = n \times M_w$$

Theoretical mass of Ag is $0.154 \times 108 = 16.63$

The yield is $\text{theor. } m / \text{real } m. \times 100\% = 15.2 / 16.63 \times 100\% = \mathbf{91,40\%}$