Henry's law can be put into mathematical terms (at constant temperature) as

$$p = k_{\mathrm{H}}\, c$$

where p is the partial pressure of the solute in the gas above the solution, c is the concentration of the solute and k_{H} is a constant with the dimensions of pressure divided by concentration.

So if partial pressure of oxygen is (20% / 100%) = 0.2 and for nitrogen it is 0.8,the concentrations of it can be found:

C for O_2 is p/K=0.2 / 2 000 000 kPa = 0.000 000 1 M

C for N_2 is p/K=0.8 / 5 000 000 kPa = 0,000 000 16 M

But for finding composition of final gas mixture you need to know mass of water sample or volume of starting gas mixture.

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