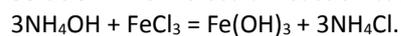
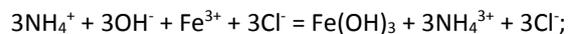


How do i find the net ionic equation for the reaction between ammonium hydroxide and iron(III)chloride ?

Solution. The molecular reaction can be described with the following equation:



Net ionic equation describes the reaction between the ions present on the solution. Ammonium hydroxide and iron(III)chloride dissociate and give the next ions: NH_4^+ , OH^- , Fe^{3+} , Cl^- . $\text{Fe}(\text{OH})_3$ doesn't dissociate to give the ions, because it is a solid compound, it is a precipitate. NH_4Cl is a soluble in the water salt. It gives NH_4^+ and Cl^- ions. The molecular reaction can be overwritten into total ionic equation:



You can see, that in the in both left and right parts there are the same ions: NH_4^+ and Cl^- and they can be reduced. In this case the net ionic equation is formed:

