

Problem: molecular mass of acetic acid in benzene.

Solution: No matter in what solvent the substance will be diluted. The molar mass is a constant for different molecules. The molecule of acetic acid consists of Carbon, Oxygen and Hydrogen. Formula can be described as: CH_3COOH .

$$M(\text{CH}_3\text{COOH}) = 2 * M(\text{C}) + 4 * M(\text{H}) + 2 * M(\text{O}) = 60 \text{ g/mol}$$

Answer: $M(\text{CH}_3\text{COOH}) = 60 \text{ g/mol}$