

28062, Chemistry, Inorganic Chemistry

A student weighs 2.046 g Cu into a 250 mL beaker at the beginning of the experiment. At the end of the experiment 1.347 g Cu is recovered. Calculate the percent loss of Cu?

Solution:

The loss of Cu at the experiment is: $2.046 - 1.347 = 0.699$ g.

So, the percent loss of Cu is: $\frac{0.699}{2.046} 100\% = 34.16\%$.

Answer: the percent loss of Cu is 34.16%.