

Chloramine are derivatives of ammonia by substitution of one, two or three hydrogen atoms with chlorine atoms. Monochloramine is an inorganic compound with the formula NH_2Cl .

sp^3 hybridization

geometry: tetrahedral

OF_2 :

Oxygen is the central atom and there will be single bonds to the fluorine's. The oxygen atom will have two lone pairs and the fluorine's will each have three lone pairs. The geometry is bent because of the lone pairs.

Actually the molecule would look very much like water.

The hybridization of all the molecules is sp^3 .

The bond should be a polar covalent bond because the electronegativity difference of the O --- F bond is greater than 0.4.

Oxygen will have a partial positive charge and the fluorines would have a partial negative charge because the fluorine is more electronegative so it will pull electrons away from oxygen.